

Name _____ Mod _____ Student Number _____

Lab # _____ **SILVER**

Purpose: In this laboratory exercise you will: (a) make note of the physical characteristics of the element silver, (b) observe a single replacement reaction in which copper is used to free atoms of silver from a solution of silver nitrate, (c) note a characteristic change in the color of the solution as more copper ions (Cu^{+2}) become present

Materials: copper wire (Cu), silver nitrate solution (AgNO_3), 50 ml beaker, glass tubing, dropping pipette, plastic gloves

Procedure: SILVER - physical properties (Complete the following on this sheet of paper)

- | | |
|----------------------|-------------------------------|
| a. symbol | j. number of neutrons |
| b. atomic number | k. number of electrons |
| c. atomic weight | l. valence |
| d. state | m. electron dot formula |
| e. color code | n. specific gravity (density) |
| f. metal/non-metal | o. color code (S. Welch) |
| g. period | p. atomic diagram |
| h. group | |
| i. number of protons | |

Replacing the silver -

